

(19)



Europäisches Patentamt  
Europ an Patent Office  
Office uropéen des brev ts

(11) Publication number:

**0 153 841  
A2**

(12)

## EUROPEAN PATENT APPLICATION

(21) Application number: 85301080.9

(51) Int. Cl.<sup>4</sup>: C 01 B 11/02

(22) Date of filing: 19.02.85

(30) Priority: 22.02.84 CA 447983

(43) Date of publication of application:  
04.09.85 Bulletin 85/36

(84) Designated Contracting States:  
AT BE DE FR GB IT NL SE

(71) Applicant: Tenneco Canada Inc.(ERCO DIVISION)  
2 Gibbs Road  
Islington Ontario(CA)

(72) Inventor: Rapson, William Howard  
110 Bloor Street West Suite 1001  
Toronto Ontario(CA)

(72) Inventor: Fredette, Maurice Conrad Joseph  
1316 Willowbank Trail  
Mississauga Ontario(CA)

(74) Representative: Hamilton, Raymond et al,  
c/o Albright & Wilson Limited 1 Knightsbridge Green  
London SW1X 7QD(GB)

(54) Production of chlorine dioxide.

(57) Chlorine dioxide, particularly for water treatment but useful in other applications of chlorine dioxide, such as the bleaching of pulp, is produced in an enclosed reaction zone filled with chlorine dioxide-generating reaction medium under sufficient pressure to prevent chlorine dioxide from forming a continuous gaseous phase and the resulting aqueous solution of chlorine dioxide in spent reaction medium is discharged to a recipient aqueous medium, such as a flowing water body to be treated, without the formation of gaseous phase chlorine dioxide.

EP 0 153 841 A2

The present invention relates to the production of chlorine dioxide and the use thereof in the treatment of water, bleaching of pulp or other applications wherein chlorine dioxide is useful.

The use of chlorine dioxide in the treatment of water for disinfection and removal of oxidizable species, such as carbonaceous and sulphurous materials, is well known and is known to have advantages over the use of chlorine. For example, chlorine dioxide avoids the formation of chlorinated end products which are mutagenic and chlorine dioxide is capable of eliminating chlorinated phenolic tastes. Relatively small quantities of chlorine dioxide are required in such use, typically 250 to 1000 kg/day in a municipal sewage treatment plant of a medium-sized community.

Usually, chlorine dioxide is formed in a reaction zone from chlorine dioxide-generating reactants in an aqueous acid reaction medium and is removed from the reaction zone in a gaseous product stream in which the partial pressure of chlorine dioxide is maintained sufficiently low to prevent decomposition of the chlorine dioxide. The chlorine dioxide gas then is dissolved in water for later use or directly in the water to be treated by the chlorine dioxide. In addition to producing gaseous products of reaction, which must be processed to an aqueous solution form, previous efforts at producing chlorine dioxide in a small compact unit with high yield, short residence time and high production rates have failed, mainly as a result of slow or incomplete reaction.

In accordance with the present invention, the chlorine dioxide-producing reaction medium fills the reaction zone and sufficient superatmospheric pressure

0153841

is applied to maintain potentially gaseous products of reaction, including chlorine dioxide, dissolved in the aqueous phase, so that no continuous gaseous product phase is formed. Upon discharge of the aqueous solution of chlorine dioxide in spent reaction medium to recipient aqueous medium, such as a flowing body of water, the chlorine dioxide remains dissolved in the aqueous phase. It is unnecessary, therefore, to provide a diluent gas and gaseous phase handling and processing equipment. While the invention has particular application to the treatment of water by chlorine dioxide and is specifically described with reference thereto, the chlorine dioxide produced using the process of the invention may be used for other known applications of chlorine dioxide including the bleaching of pulp.

By utilizing this process, it is possible to achieve high yield and high production rate and thereby overcome the prior art problems noted above, particularly when the chlorine dioxide-generating reactants are sodium chlorate, sodium chloride and sulphuric acid. It has been found, in accordance with one particularly-preferred embodiment of the invention, that substantially all the sodium chlorate can be reacted at a high yield in excess of 85% by using reactant feed rates which result in a reaction medium containing about 0.8 to 1 molar sodium chlorate, about 0.8 to 1 molar sodium chloride and about 15 to 16 normal sulphuric acid, a residence time of about 25 to about 150 seconds or longer and a pressure of about 400 to about 550 kilo pascal (kPa) to maintain the chlorine dioxide and chlorine dissolved in the aqueous phase. The reactions which occur in the reaction zone and the enclosed nature of the equipment which is employed do not permit direct measurement of the actual concentrations of reactants in the reaction medium. The concentration values for the reactants

recited in this paragraph and elsewhere in this disclosure are calculated from feed concentrations after mixing and before reaction, and from product yield.

5 In accordance with another preferred embodiment of the invention, chlorine dioxide is formed from a reaction medium containing about 0.8 to 1 molar sodium chlorate and about 4 to 6 normal hydrochloric acid, a residence time of about 20 to about 200 seconds or  
10 longer and a pressure of about 400 to about 550 kPa to maintain the chlorine dioxide and chlorine dissolved in the aqueous phase.

In a yet further preferred embodiment of the invention, the reaction medium contains about 0.8 to  
15 about 1 molar sodium chlorite and about 0.01 to 0.1 normal hydrochloric acid, a residence time of about 20 to about 200 seconds or longer and a pressure of about 400 to about 550 kPa.

In the present invention, the aqueous solution of  
20 chlorine dioxide in spent reaction products can be conveniently discharged directly into a body of water to be treated, or otherwise employed. This solution contains residual sulphuric acid and sodium salts, but the concentrations are insufficient to cause concern  
25 in water treatment for potable use. However, it is important to achieve substantial decomposition of the sodium chlorate reactant before discharge, in view of the known toxicological problems of this substance in water.

30 In the process of the present invention, chlorine dioxide-producing reactants are fed to an enclosed reaction zone. At least one of the reactants is fed to the reaction zone in the form of an aqueous solution thereof, to provide an aqueous reaction  
35 medium filling the reaction zone.

One reactant should be sodium chlorate, sodium chlorite or both. The sodium chlorate and/or sodium

chlorite is reacted with at least one other reactant fed to the reaction zone to form chlorine dioxide, with or without chlorine. In general, one reactant may contain sodium chlorate and any suitable reducing agent while the other reactant may be any strong acid, such as, hydrochloric acid, sulphuric acid or phosphoric acid. Combinations of suitable reactants which may be used in this invention include sodium chlorate, sodium chloride and sulphuric acid; sodium chlorate and hydrochloric acid; sodium chlorate, sodium chlorite, sodium chloride and sulphuric acid; sodium chlorate, sodium chlorite and hydrochloric acid; sodium chlorate, sulphur dioxide and sulphuric acid; sodium chlorate, methanol and sulphuric acid; sodium chlorite and chlorine, in gaseous or aqueous solution form; sodium chlorite and hydrochloric acid, sulphuric acid or other strong acid; sodium chlorite, suitable oxidizing agent and sulphuric acid; sodium chlorate, sodium chloride, hydrogen peroxide and/or methanol and sulphuric acid; sodium chlorite, sodium hypochlorite and hydrochloric acid and/or sulphuric acid; and sodium chlorate, glucose and sulphuric acid. Phosphoric acid may be used as partial or complete substitute for sulphuric acid.

In many chlorine dioxide-generating processes, chlorine is coproduced with chlorine dioxide. If desired, the chlorine may be reduced to chloride form by addition of a suitable reagent, such as, hydrogen peroxide, to the reaction product solution before use thereof.

The reactants may be fed to the reaction zone in any convenient manner to achieve rapid intermixing of the reactants. In one embodiment of the invention, a first aqueous reactant stream is fed under pressure in a pipeline conveyor to and through a venturi or constricted throat. A second reactant stream enters the centre of the venturi under the influence of the

0153841

suction induced by the flow of the first aqueous reactant through the venturi. The second reactant stream may be a further aqueous solution, or may be a gaseous reactant. The turbulence in the venturi throat mixes the reactants thoroughly and rapidly. Under the influence of the pressure in the pipeline, the chlorine dioxide and other potentially gaseous reaction products, including chlorine, remain in the aqueous phase. The resulting aqueous solution of chlorine dioxide in spent reaction medium may conveniently be forwarded to a rapidly flowing, much larger stream of water to be treated, the chlorine dioxide remaining dissolved in the aqueous phase and disinfecting or otherwise treating the water stream. The aqueous chlorine dioxide solution in spent reaction medium may be discharged to the water stream being treated through a back-pressure control valve which prevents the flow of water back into the chlorine dioxide solution stream.

In another embodiment of the invention, the reactant streams are fed to a feed tee located at one end of a pipe reactor. The feed tee is equipped with inlet nozzles which promote mixing of the reactants prior to entering the pipe reactor. The reactants are maintained under pressure in the pipe reactor, so that gaseous chlorine dioxide, and possibly gaseous chlorine depending on the reactants, remains dissolved in the aqueous phase and a continuous gaseous phase is not formed. At the other end of the reactor, the aqueous solution of chlorine dioxide may be diluted before discharge to a storage tank at atmospheric pressure for later use.

Since the formation of a continuous gaseous chlorine dioxide phase is avoided in this invention, there is no necessity for the provision of a source of diluent gas to maintain the partial pressure of gaseous chlorine dioxide below decomposition values.

Further, there is no requirement for separate hardware to effect dissolution of the gaseous chlorine dioxide for its subsequent use. In the present invention, the pressurized aqueous chlorine dioxide solution which results from the reaction should be used in a manner which avoids the formation of an undiluted gas phase to prevent explosion.

The equipment which is required in this invention is simple in construction and requires little capital. The process is ideally suited to produce chlorine dioxide in quantities suitable for disinfection and other treatment of water.

The invention is described further, by way of illustration, with reference to the accompanying drawings, in which:

Figure 1 is a schematic representation of a flow sheet for a process of water treatment using the present invention;

Figure 2 is a longitudinal sectional view of a chlorine dioxide producing apparatus utilized in the process of Figure 1; and

Figure 3 is a sectional view taken on line 3-3 of Figure 2.

Referring to the drawings, a water treatment operation utilizes a chlorine dioxide-producing apparatus 12, shown schematically in Figure 1 and described in more detail below with respect to Figures 2 and 3, which receives reactant solutions by lines 14 and 16 and which discharges chlorine dioxide in aqueous solution form to a body of water 18 being treated and flowing through pipe 20.

In the specific process illustrated in Figure 1, the reactant solution in line 14 is an aqueous solution of sodium chlorate and sodium chloride while the reactant solution in line 16 is sulphuric acid. It will be apparent from the above discussion of the

0153841

invention that other combinations of reactants may be used.

5 The aqueous solution of sodium chlorate and sodium chloride in line 14 is pumped to the generator 12 by a suitable metering pump 22 communicating with a storage tank 24 containing the solution. The concentration of sodium chlorate and sodium chloride in the aqueous feed solution may vary from about 1 to about 2.5 molar each, typically about 2 molar, usually 10 in molar proportions of sodium chlorate to sodium chloride of about 0.9:1 to about 0.95:1, corresponding to a chloride ion to chlorate ion ratio of about 1:1.05.

15 The sulphuric acid in line 16 is pumped to the generator 12 by a suitable metering pump 26 communicating with a storage tank 28. The sulphuric acid feed solution usually has a high titratable acid normality in the range of about 20 to about 36 normal, typically about 30 normal.

20 The aqueous solution of sodium chlorate and sodium chloride and the sulphuric acid are pumped to the generator 12 to provide, in the generator 12, a highly acidic reaction medium which is capable of rapid generation of chlorine dioxide and chlorine. 25 The titratable acid normality of the reaction medium after mixing usually is in the range from about 8 to about 16 normal, preferably about 15 to 16 normal. The calculated concentration of sodium chlorate and sodium chloride in the highly acidic reaction medium 30 usually vary from about 0.5 to about 1.2 molar each, preferably about 0.8 to about 1 molar.

35 As a result of the highly concentrated nature of the sulphuric acid which is fed to the generator 12 and the rapid dilution thereof upon mixing with the sodium chlorate and sodium chloride solution, the heat of dilution generated thereby results in a reaction medium usually having a temperature in the range of



0153841  
about 30° to about 50°C., preferably about 35° to about 45°C. It will be obvious to those skilled in the art that higher temperatures may be used if corresponding higher pressures are used to maintain the chlorine dioxide dissolved.

5 The highly acid reaction medium fills the generator 12 and is subjected to a superatmospheric pressure therein to substantially prevent the formation of a continuous phase of gaseous chlorine dioxide and chlorine and substantially retain those  
10 gases in the dissolved state in the reaction medium. Usually, the pressure is maintained at about 400 to about 620 kPa, preferably about 480 to about 550 kPa, in the generator 12.

15 The pressurized solution of chlorine dioxide and chlorine in spent reaction medium is discharged from the generator to the water 18 flowing in pipe 20 after a residence time at least sufficient to effect substantially complete reaction of the sodium  
20 chlorate, usually about 12 to about 300 seconds, preferably about 25 to about 150 seconds, although longer residence times may be employed, if desired.

The discharge of the chlorine dioxide and chlorine solution into the body of water 18 achieves  
25 substantial dilution of the chlorine dioxide and chlorine, so that the gases remain dissolved and effect the desired treatment of the water in the pipe 20. Usually, the chlorine dioxide-containing solution is diluted about 10 to about 30 times, typically about  
30 twenty times.

Details of construction of a suitable chlorine dioxide generator 12 are shown in Figures 2 and 3, to which reference now is made. The chlorine dioxide generator comprises a pair of concentrically-arranged  
35 tubes 30 and 32. The inner tube 30 is constructed of corrosion-resistant material, typically a corrosion-resistant polymeric material, and the outer

0153841

tube 32 is constructed of rigid deformation-resistant material, such as stainless steel or similar metal.

5 The outer rigid tube 32 protects the inner tube 30 from deformation as a result of the flow of the water stream 18 and of the generation of heat of reaction and acid dilution, and hence stabilizes the structure. The corrosion-resistant nature of the inner tube 30 results in a long life of operation in the highly acid environment. If desired, a metal pipe  
10 lined with corrosion-resistant material, such as glass, may be used in place of the concentric tubes 30 and 32.

The inner tube 30 defines a reaction chamber 34 for the chlorine dioxide-producing reaction and  
15 extends from an annular flange 36 to an outlet member 38 comprising a conical surface 40 converging to a narrow outlet bore 42. Tubes of differing lengths and/or diameters may be used to provide reaction chambers of varying capacity.

20 In place of the fixed orifice construction of the outlet bore 42, there may be utilized a back-pressure control valve which permits adjustment of the back pressure within the reaction chamber 34 for different production rates. In addition, a pressure operated  
25 valve which opens when a predetermined pressure, which may be variably predetermined, is reached in the reaction chamber 34, also may be used as the outlet from the generator 12.

The annular flange 36 and the outlet member 38  
30 are constructed of corrosion-resistant material, typically the same material of construction as the tube 30. The annular flange 36 is sandwiched between an outer closure disc 44 of corrosion-resistant material and a further annular flange 46 to which is  
35 affixed one end of the outer tube 32, annular sealing gaskets 48 being provided between abutting surfaces.

0153841

A reactant mixing chamber 50 is defined between the central bore 52 of the flange 36 and the end closure disc 44. Two reactant inlet bores 54 and 56 communicate with the reactant feed lines 14 and 16. One inlet bore 54 for sodium chlorate and sodium chloride solution extends through the flange 36 from the external surface thereof to the mixing chamber 50, exiting to the mixing chamber 50 substantially tangential thereto, while the other inlet bore 56 for sulphuric acid extends through the end closure disc 44 axially of the mixing chamber 50, to promote rapid intermixing of reactants fed by the inlet bores 54 and 56 which communicate with the reactant feed lines 14 and 16.

An alternative structure of mixing chamber for the generator 12 involves the use of a venturi which would have one of the reactant feeds, preferably the sodium chlorate and sodium chloride solution, communicating with the convergent pipe of the venturi and the other of the reactant feeds, preferably the sulphuric acid, communicating with the collar of the venturi. The flow of the one reactant then would draw the other reactant into the collar for rapid and intimate mixing.

As mentioned previously, the process of the invention is not limited to the specific combination of reactant feeds mentioned in connection with lines 14 and 16 but rather the invention has general application to any combination of chlorine dioxide-generating reactants. Specific alternate combinations are discussed above.

The outer tube 32 extends from the flange 46 to a cup-like end member 58 which sealingly engages the outer surface of the outlet member 38 against the ingress of the water from the flowing stream 18 to between the inner and outer tubes 30 and 32. The cup-like end member 58 has a central opening 60

through which projects the cylindrical portion 62 of the outlet 38 having the bore 42 therethrough.

5 The generator 12 is mounted to an elbow 64 of the pipe 20 with the concentric tubes 30 and 32 extending approximately vertically upwardly into the flowing water stream 18 and terminating in a vertical portion of the pipe 20. At the exit of the generator 12, therefore, the flow of chlorine dioxide solution and water to be treated is in the same direction.

10 While the generator 12 is illustrated as discharging the chlorine dioxide through an elbow in the pipe 20 and in the direction of flow of the water, this arrangement is provided for convenience of assembly and is not critical. The generator 12 may be  
15 arranged substantially radially of a straight pipe-section or any other convenient arrangement.

The generator 12 is mounted to the pipe elbow 64 by bolting to an annular flange 68 at the end of a vertical pipe extension 70, with an annular gasket 71  
20 being located between the abutting surfaces of the flanges 46 and 68.

A tangential bore 72 is through the flange 36 in communication with a valved pressurized wash water feed line 74. Normally, the valve in the water feed  
25 line 74 is closed, but upon ceasing of chlorine dioxide production as a result of the termination of the flow of reactants to the generator 12, the valve is opened to permit wash water to flow into the mixing chamber 50 and the reaction chamber 34 to flush out  
30 spent reaction medium.

In operation of the apparatus of Figures 1 to 3, an aqueous solution of sodium chlorate and sodium chloride enters the generator 12 through bore 54 while sulphuric acid enters the generator 12 through bore  
35 56. The reactants rapidly intermix in the mixing chamber 50 and flow through the reaction chamber 34 inside the pipe 30 under the influence of the pressure

0153841

of flow of reactants to the mixing chamber 50. The  
reducing diameter of the pipe 30 at the cone 40  
terminating in the restricted diameter outlet bore 42,  
or valved outlet, serves to apply back pressure on the  
5 reaction medium in the pipe 30, so as to prevent the  
formation of a continuous gaseous phase of chlorine  
dioxide and chlorine and retain those  
potentially-gaseous products of reaction substantially  
dissolved in the reaction medium. The resulting  
10 aqueous solution of chlorine dioxide and chlorine  
dissolved in spent reaction medium is discharged  
through the outlet bore 42 to the flowing water stream  
18 and is dissolved and dispersed therein.

The generator 12 is simple in construction, easy  
15 to operate and requires no maintenance. Chlorine  
dioxide can be produced rapidly at high yield and good  
efficiency as an aqueous solution in spent reaction  
medium. No gaseous diluents are required and the  
equipment required for handling gaseous reaction  
20 products is eliminated.

The invention is illustrated further by the  
following Examples:

Example 1

An experimental apparatus was set up comprising a  
25 Y-joint communicating with the upper end of a  
capillary tube 140 cm long and 2.5 to 3 mm in  
diameter. The first 10 cm of the leg of the Y-joint  
contained constrictions to promote mixing of the  
reactant streams fed to the arms of the Y-joint. A  
30 cooling jacket was provided surrounding the capillary  
tube. The lower exit of the capillary tube  
communicated with a tee to which water was fed to  
dilute the reaction products and remove the diluted  
solution to a collection jar.

35 An aqueous solution of sodium chlorate and sodium  
chloride, pressurized to about 480 kPa, was fed to one  
inlet of the Y-joint while sulphuric acid, also

0153841

pressured to about 480 kPa, was fed to the other inlet of the Y-joint. The reaction medium filled the tube and the formation of a continuous gaseous phase was not observed. A range of concentrations of sodium chlorate, sodium chloride and sulphuric acid, with varying residence times, was tested to attempt to establish conditions which result in reaction of substantially all the chlorate using the least amount of acid. The reactor was cooled to maintain a reaction temperature of about 25° to 30°C.

The results obtained are reproduced in the following Table I:

TABLE I

Run	1 (4)	2 (4)	3	4	5	6	7	8	9	10	11	12	13
Feed $\text{ClO}_3^-$ (M)	3	3	3	3	2	2.3	3	3.4	2	2.1	2	2	1.6
$\text{Cl}^-$ (M)	3	3	3	3	2	2.3	3	3.4	2	2.1	2.3	2.3	1.7
$\text{ClO}_3^-$ in reactor (M)	N.D. (1)	N.D.	N.D.	N.D.	1.03	1.0	1.75	1.35	1.05	0.9	0.88	1.0	0.8
$\text{H}^+$ in reaction (N)	9.4	9.4	12.9	13.6	13.6	15.7	12.1	17	13.5	16	15.8	14.1	14
% $\text{ClO}_3^-$ reacted	56	75	71	70	64	96.6	63	58	76	98	74	80	79
% Efficiency (2)	68	30	87.3	80.7	83.6	76	59	74	81	87.6	64	95	78
% Yield (3)	30	23	62	50.5	53.7	74	37	43	61	85.7	47	76	62
Residence Time (secs)	12	20	22	47	27	25	20	26	22	26	11	21	23

Notes: (1) N.D. means not determined

(2) Efficiency =  $\frac{\text{ClO}_2 \text{ formed}}{\text{ClO}_3^- \text{ reacted}}$ (3) Yield =  $\frac{\text{ClO}_2 \text{ formed}}{\text{Total ClO}_3^- \text{ feed}}$ 

(4) No cooling

0153841

It will be seen from the data reproduced in the above Table I that, in the first four runs at a feed of 3 molar sodium chlorate and sodium chloride, increasing the acidity in the reaction medium from 9.4 to 13.6 normal and increasing the residence time from 12 to 42 seconds had little effect on the approximately 70% maximum reaction of sodium chlorate. However, by lowering the concentration of the feed solution of sodium chlorate and sodium chloride, so as to provide a chlorate concentration in the reaction medium of about 0.9 to 1 molar, and by increasing the acidity of the reaction medium to about 15 to 16 normal, 96.6% and 98% conversion of sodium chlorate was achieved (runs 6 and 10 respectively). In the case of run 10, the substantially complete reaction of sodium chlorate was achieved at an efficiency of 88% and a yield of 86%.

#### Example 2

A further experimental apparatus was set up comprising a tee communicating with the lower end of a glass tube 30 cm long with an inside diameter of 2.54 cm. Nozzles were installed in the tee to inject reactant feeds into the reactor. A second tee was provided at the upper end of the tube with a diluent water feed line communicating with the tee to contact the reaction products and dilute the same. The exit of the upper tee communicated with a product collection vessel.

An aqueous solution of sodium chlorate and sodium chloride was fed to one inlet to the lower tee while sulphuric acid was fed to the other inlet to the lower tee. In some runs, hydrogen peroxide or methanol also was fed with the sodium chlorate and sodium chloride solution to decrease the quantity of chlorine coproduced with the chlorine dioxide. The flow rates of reactants were such as to produce a pressure in the reaction tube of 450 to 515 kPa. In each case,



0153841

reaction medium filled the tube and the formation of a continuous gaseous phase was not observed.

The results are reproduced in the following Table II:

5

10

15

20

25

30

35

TABLE II

Run NO.	F1 (M)		H <sub>2</sub> O <sub>2</sub>		F2 (N)		Feed Rate		Pressure (kPa)	AE	%Y	H <sup>+</sup> (N)	%ClO <sub>3</sub> <sup>-</sup> Reacted
	ClO <sub>3</sub> <sup>-</sup>	Cl <sup>-</sup>	H <sub>2</sub> O <sub>2</sub>	MeOH	F2 (N)	H <sub>2</sub> SO <sub>4</sub>	F1	F2					
1.	2.06	2.34			30		28.9	37.0	450	82	71.7	16.8	88
2.	2.04	2.02			30		40.6	19.4	575	73	33.0	9.7	45
3.	3.90	1.00	1.95		33		31.2	25.3	480	96	24.9	14.8	26
4.	2.10	0.54	1.05		33		27.4	26.7	480	71	35.6	16.3	50
5.	1.86	2.10	1.05		33		20.1	28.3	480	83	53.3	19.3	65
6.	2.10	2.10			30		29.4	33.2	480	82	71.0	16.0	87
7.	1.80	1.83			30		28.4	29.4	480	94	85.1	15.2	96
8.	1.77	1.00	0.90		33		27.9	29.3	480	79	65.2	15.4	83
9.	1.45	0.96	0.90		33		21.1	23.5	480	71	64.0	17.4	91
10.	1.26	1.24	0.62		33		23.3	23.6	480	70	59.5	16.6	96
11.	1.83	0.03		0.72	33		26.1	23.0	480	-	2.0	15.5	-
12.	1.86	1.90		0.72	33		22.5	22.1	480	58	49.1	16.5	85

0153841

The results of the above Table II generally corroborate those of Table I for feeds of sodium chlorate and sodium chloride. In run 7, 94% efficiency was observed with 96% reaction of sodium chlorate. In runs where hydrogen peroxide was added to the feed yields ranged from 25 to 65%, while methanol yielded similar results.

In another experiment, hydrogen peroxide was added to the product solution resulting from reaction of sodium chlorate, sodium chloride and sulphuric acid in the reactor. It was found that the hydrogen peroxide preferentially reduced the chlorine with little effect on  $\text{ClO}_2$ , yielding a solution with a  $\text{ClO}_2/\text{Cl}_2$  ratio of 19:1 w/w, with overall yield near 84%.

### Example 3

Runs were conducted using the apparatus described in Example 2 for reactant streams comprising sulphuric acid and an aqueous solution of sodium chlorate, sodium chlorite and sodium chloride as the feeds. The reaction medium filled the reactor tube and the formation of a continuous gaseous phase was not observed. The results are set forth in the following Table III:

TABLE III

Run No.	Feed Composition (M) NaClO <sub>3</sub> NaClO <sub>2</sub> NaCl	Feed Rate (moles/hr) NaClO <sub>3</sub> NaClO <sub>2</sub> NaCl	H <sup>+</sup> (N) H <sub>2</sub> SO <sub>4</sub>	Yield	%ClO <sub>3</sub> <sup>-</sup> Reacted
1	1.09 1.01 1.05	1.68 1.56 1.62	46.3 15	85	100
2	1.11 1.04 1.04	1.25 1.25 1.33	36.2 15	88.3	92.5

0153841

As may be seen from the results of the above Table III, moderately good yields can be obtained with substantial reaction of sodium chlorate.

In summary of this disclosure, the present invention provides a novel process for forming chlorine dioxide wherein the chlorine dioxide does not form a discrete gaseous phase but rather is retained in solution from formation to use. Modifications are possible within the scope of this invention.

5

10

15

20

25

30

35

1. A process for the production of an aqueous solution of chlorine dioxide, characterized in that chlorine dioxide-producing reactants are fed to a reaction zone which is filled with an aqueous reaction medium; the aqueous reaction medium is subjected to a superatmospheric pressure in the reaction zone which is sufficient to maintain potentially gaseous products of the reaction including chlorine dioxide in solution and the resulting aqueous solution of chlorine dioxide in the spent reaction medium is discharged from the reaction zone to an aqueous medium under such conditions that substantially no chlorine dioxide gas is formed.
2. A process according to claim 1, characterized in that the solution of chlorine dioxide from the reaction zone is discharged directly to an aqueous medium in order to disinfect that medium.
3. The process claimed in either of claims 1 or 2, characterized in that the reaction zone is an elongate tubular reaction zone, the chlorine dioxide-producing reactants are fed as two separate reactant streams to one end of the tubular reaction zone to be intermixed thereat, the intermixed reactants flow through the tubular reaction zone, and the aqueous solution of chlorine dioxide is discharged from the other end of the tubular reaction zone.
4. The process claimed in claim 3, characterized in that at least one of the reactant streams is fed generally tangentially to a circular mixing zone at the one end of the tubular reaction zone.
5. The process claimed in claim 3, characterized in that one of the reactant streams is fed to the converging portion of a venturi zone and the other of

the reactant streams is fed to to the collar portion of the venturi zone.

6. The process claimed in claim 3, 4, or 5, characterized in that one of the reactants comprises an aqueous solution of sodium chlorate, the other of the reactants comprises a strong acid, and a reducing agent for the sodium chlorate is provided by the acid or as a separate component in the aqueous solution of sodium chlorate.

7. An apparatus for the production of an aqueous solution of chlorine dioxide, characterized by an elongate reaction tube (12) having an inlet for reactants at one longitudinal extremity and an outlet (38) for the discharge of reaction products from the other longitudinal extremity, a mixing chamber (50) communicating with the inlet, and chlorine dioxide generating reactant feed inlets (54, 56) communicating with the mixing chamber (50).

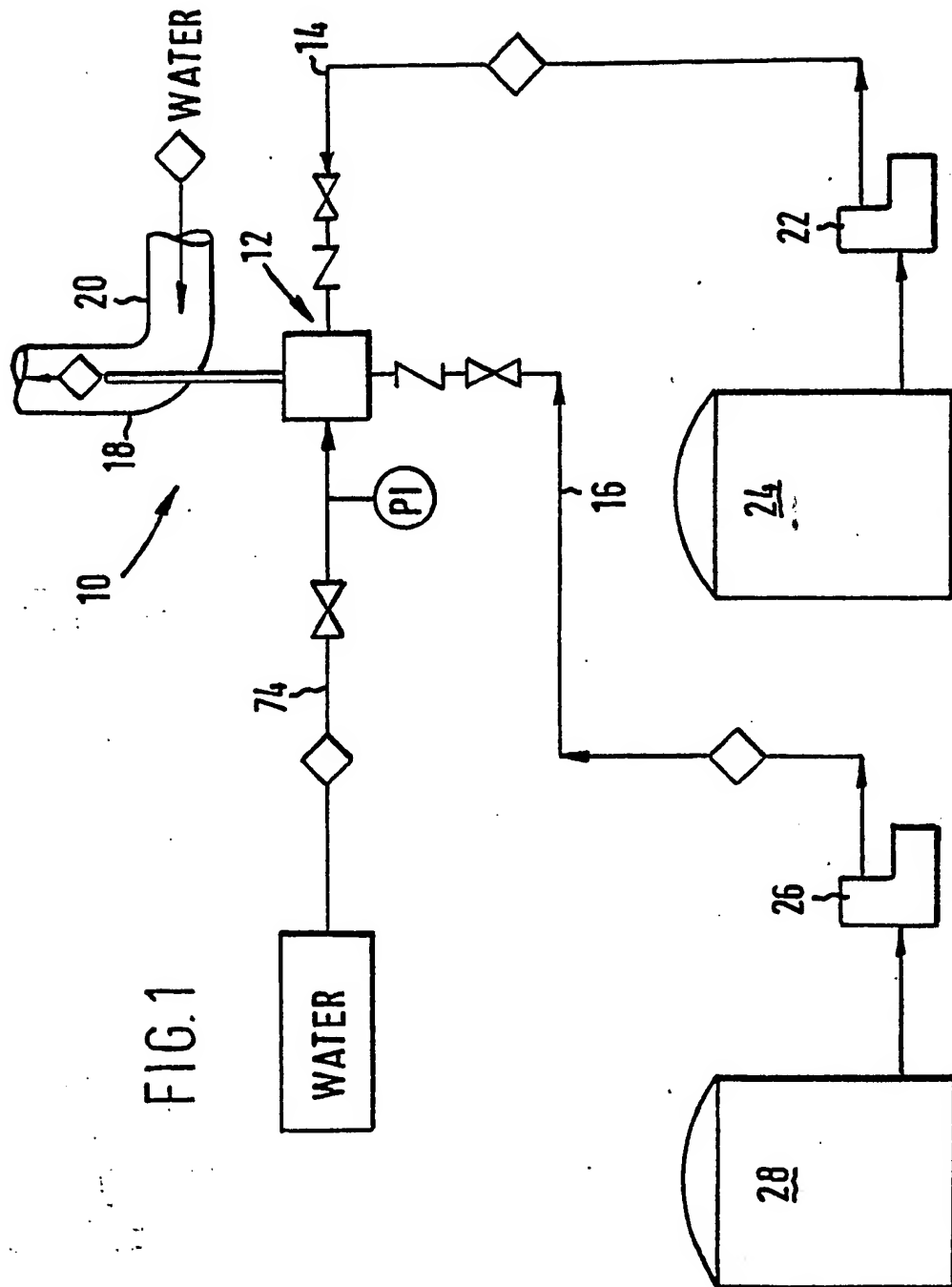
8. The apparatus claimed in claim 7, characterized in that the mixing chamber (50) is of generally circular cross-section and one of the feed inlets (54) communicates generally tangentially to the mixing chamber (50) while the other of the feed inlets (56) extends axially into the mixing chamber (50).

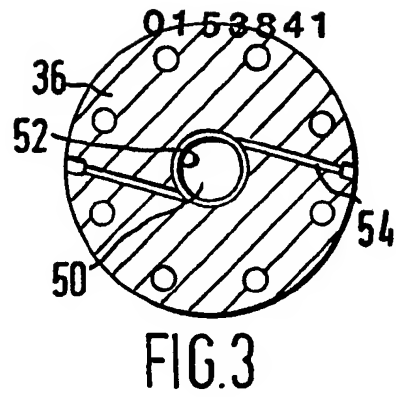
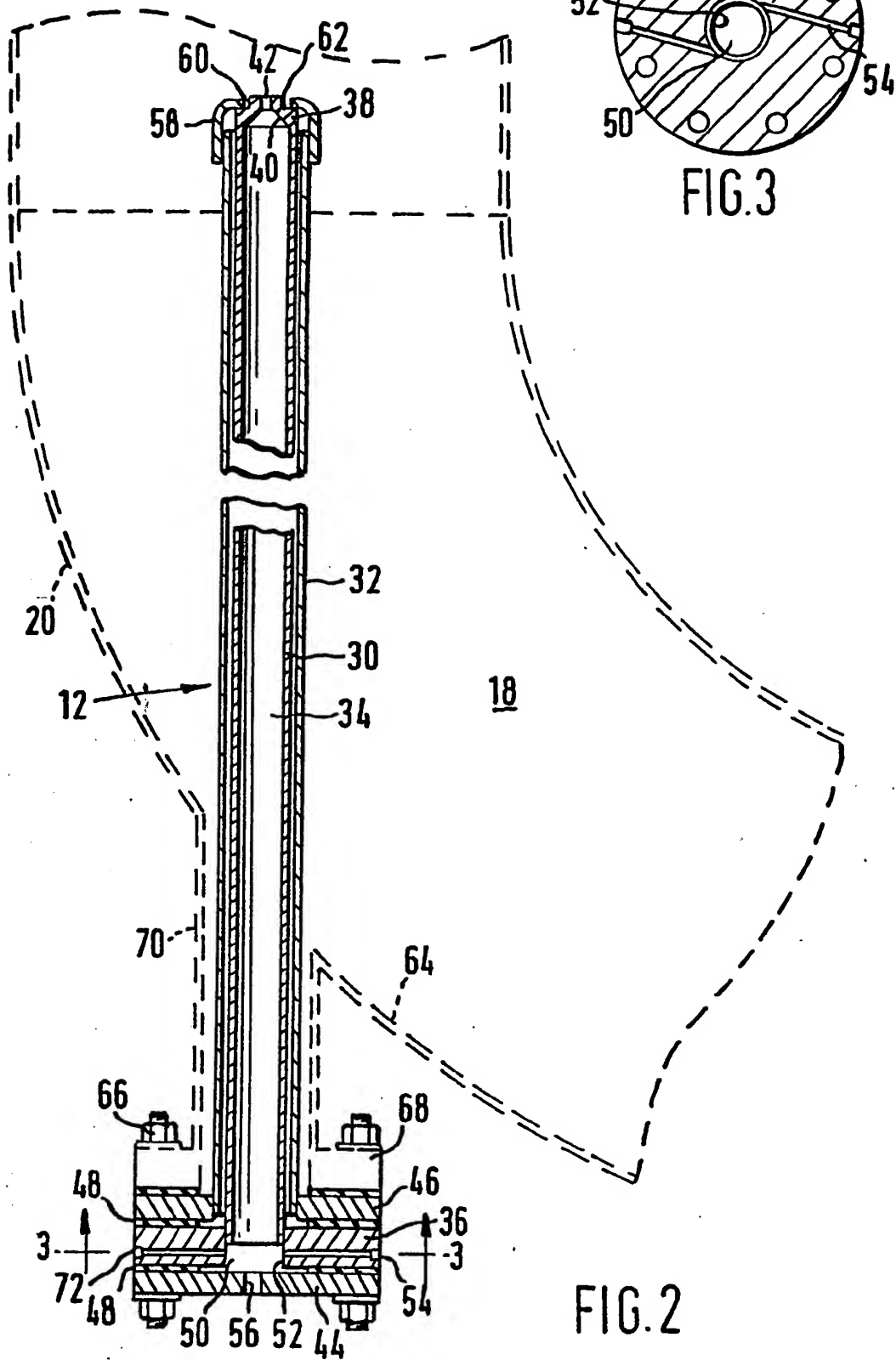
9. The apparatus claimed in claim 8, characterized in that the mixing chamber (50) is defined by an end plate (44) and a central opening in an annular flange (36) to which the reaction tube (12) is attached, the one feed inlet (54) extends in a generally radial direction through the annular flange (36), and the other feed inlet (56) extends axially through the end plate (44).

10. The apparatus claimed in claim 9, characterized in that a second bore (72) extends generally radially through the annular flange (36) to permit flushing water to be fed to the mixing chamber (50) and through the reaction tube (12) to the outlet (42).

11. The apparatus claimed in any one of claims 7 to 10 characterized in that the outlet (38) comprises a conical portion (40) of diameter decreasing in the direction of reaction medium flow and terminating in an axial bore (42) through which reaction products are discharged from the reaction tube.







(19)



**Eur päisches Patentamt**  
**European Patent Office**  
**Office eur péen des brevets**

(11) Publication number:

**0 153 841**  
**A3**

(12)

## EUROPEAN PATENT APPLICATION

(21) Application number: 85301080.9

(51) Int. Cl.<sup>4</sup>: C 01 B 11/02

(22) Date of filing: 19.02.85

(30) Priority: 22.02.84 CA 447983

(43) Date of publication of application:  
04.09.85 Bulletin 85/36

(86) Date of deferred publication of search report: 25.06.86

(84) Designated Contracting States:  
AT BE DE FR GB IT NL SE(71) Applicant: Tenneco Canada Inc.(ERCO DIVISION)  
2 Gibbs Road  
Islington Ontario(CA)(72) Inventor: Rapson, William Howard  
110 Bloor Street West Suite 1001  
Toronto Ontario(CA)(72) Inventor: Fredette, Maurice Conrad Joseph  
1316 Willowbank Trail  
Mississauga Ontario(CA)(74) Representative: Hamilton, Raymond et al,  
c/o Albright & Wilson Limited 1 Knightsbridge Green  
London SW1X 7QD(GB)

(54) Production of chlorine dioxide.

(57) Chlorine dioxide, particularly for water treatment but useful in other applications of chlorine dioxide, such as the bleaching of pulp, is produced in an enclosed reaction zone filled with chlorine dioxide-generating reaction medium under sufficient pressure to prevent chlorine dioxide from forming a continuous gaseous phase and the resulting aqueous solution of chlorine dioxide in spent reaction medium is discharged to a recipient aqueous medium, such as a flowing water body to be treated, without the formation of gaseous phase chlorine dioxide.

**EP 0 153 841 A3**



DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl. 4)
A	EP-A-0 024 851 (RIO LINDA CHEMICAL CO.)  -----		C 01 B 11/02
			TECHNICAL FIELDS SEARCHED (Int. Cl. 4)
			C 01 B 11/02
The present search report has been drawn up for all claims			
Place of search BERLIN		Date of completion of the search 19-03-1986	Examiner CLEMENT J.P.
<b>CATEGORY OF CITED DOCUMENTS</b>			
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	